

LAHORE BOARD

GRADE 10

CHEMISTRY

2018 GROUP 1

MCQ's

Section A-(MCQs)

i) Night-blindness is because of deficiency of: (Mark 1)

- A. Vitamin A
- B. Vitamin E
- C. Vitamin C
- D. Vitamin D

Answer:

- A. Vitamin A

ii) Specific heat capacity of water is : (Mark 1)

- A. $4.2 \text{ K Jg}^{-4} \text{ K}^{-1}$
- B. $4.2 \text{ Jg}^{-1} \text{ K}^{-1}$
- C. $4.4 \text{ K jg}^{-1} \text{ K}^{-1}$
- D. $4.4 \text{ Jg}^{-1} \text{ K}$

Answer:

- B. $4.2 \text{ Jg}^{-1} \text{ K}^{-1}$

iii) Which one is also called olefins: (Mark 1)

- A. Alkanes
- B. Alkenes
- C. Alkynes
- D. Alcohols

Answer:

- B. Alkenes

iv) A disease that causes bone and tooth damage is : (Mark 1)

- A. Fluorosis
- B. Cholera
- C. Jaundice
- D. Hepatitis

Answer:

A. Fluorosis

v) The colour of hydrogen iodide is:

(Mark 1)

A. Orange

B. Purple

C. Red

D. Colourless

Answer:

D. Colourless

vi) Depending upon temperature variation, atmosphere is divided into how many regions:

(Mark 1)

A. One

B. Two

C. Three

D. Four

Answer:

D. Four

vii) If $Q_c < K_c$, reaction proceeds:

(Mark 1)

A. Forward

B. Reverse

C. Equilibrium

D. Both side

Answer:

A. Forward

viii) When glucose and fructose combine they produce:

(Mark 1)

A. Starch

B. Cellulose

C. Sucrose

D. None of these

Answer:

C. Sucrose

ix) Which acid causes the acidity of stomach:

(Mark 1)

A. Sulphuric acid

B. Hydrochloric acid

C. Nitric acid

D. Oxalic acid

Answer:

B. Hydrochloric acid

x) You want to dry a gas, which one of the following salt you will see:

(Mark 1)

- A. NaCl
- B. CaCO_3
- C. CaO
- D. Na_2SiO_3

Answer:

- C. CaO

xi) Percentage of nitrogen in urea is:

(Mark 1)

- A. 36.6 %
- B. 46.6%
- C. 56.6%
- D. 66.6%

Answer:

- B. 46.6%

xii) Which one of the following is not a fossil fuel:

(Mark 1)

- A. Biogas
- B. Coal
- C. Natural gas
- D. Petroleum

Answer:

- A. Biogas

Q.2 i) what is meant by active mass? Also, write its unit.

(Marks 2)

Q.2 ii) What is meant by reversible reactions?

(Marks 2)

Q.2 iii) Write down chemical equilibrium state?

(Marks 2)

Q.2 iv) Define law of mass action .

(Marks 2)

Q.2 v) Write two important properties of salts.

(Marks 2)

Q.2 vi) Write two uses of sulphuric acid.

(Marks 2)

Q.2 vii) State Arrhenius concept of acids and bases.

(Marks 2)

Q.2 viii) Define adduct.

(Marks 2)

Q.3 i) Define carbonization .

(Marks 2)

Q.3 ii) Write the name of four different types of coal.

(Marks 2)

Q.3 iii) Define alcoholic group with one example.

(Marks 2)

Q.3 vi) What is difference between oil and ghee?

(Marks 2)

- Q.3 vii) Write the basic unit of protein, give one example . (Marks 2)
- Q.3 viii) Write the name of two diseases caused by deficiency of vitamin A. (Marks 2)
- Q.4 i) Name the major constituents of troposphere. (Marks 2)
- Q.4 ii) Write four natural systems of our earth (only names). (Marks 2)
- Q.4 iii) Why CO_2 is called a greenhouse gas? (Marks 2)
- Q.4 iv) Mention the disadvantages of detergents. (Marks 2)
- Q.4 v) How water-borne disease can be prevented? (Marks 2)
- Q.4 vi) Write down two advantages of solvay's process. (Marks 2)
- Q.4 vii) Define petroleum and crude oil. (Marks 2)
- Q.4 viii) Define residual oil with two fractions name. (Marks 2)
- Q.5 a) Define equilibrium constant. How the extent of reaction can be predicted with the help of value of K_c ? (Marks 5)
- Q.5 b) Write down the four uses of bases. (Marks 4)
- Q.6 a) Write down five physical properties of alkenes. (Marks 5)
- Q.6 b) Write a note on Deoxyribonucleic acid (DNA). (Marks 4)
- Q.7 a) What is meant by concentration of Ore? Also give two methods of concentration of Ores. (Marks 5)
- Q.7 b) Write two methods for removal of permanent hardness of water. (Marks 4)

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CHEMISTRY

2018 GROUP 2

MCQ's

Section A-(MCQs)

i) General formula of carbohydrates is: (Mark 1)

- A. $C_n(H_2O)_n$
- B. $C_n(H_2O_2)_n$
- C. $C_n(H_3O)_n$
- D. $C_n(H_2O_3)_n$

Answer:

- A. $C_n(H_2O)_n$

ii) The density of water at 4°C is: (Mark 1)

- A. 1 gm cm^{-3}
- B. 2 gm cm^{-3}
- C. 3 gm cm^{-3}
- D. 4 gm cm^{-3}

Answer:

- A. 1 gm cm^{-3}

iii) Dehydration of alcohols can be carried out with: (Mark 1)

- A. H_2SO_4
- B. HCl
- C. KOH
- D. NaOH

Answer:

- A. H_2SO_4

iv) Which one of the following gases is used to destroy harmful bacteria in water: (Mark 1)

- A. Bromine
- B. Chlorine
- C. Fluorine
- D. Iodine

Answer:

B. Chlorine

v) The unit of molar concentration is:

(Mark 1)

A. mol dm⁻¹

B. mol dm⁻²

C. mol dm³

D. mol dm⁻³

Answer:

D. mol dm⁻³

vi) About 99 % of atmospheric mass lies is:

(Mark 1)

A. 16 km

B. 17 km

C. 30 km

D. 35 km

Answer:

C. 30 km

vii) In the beginning, the rate of reverse reaction is:

(Mark 1)

A. Moderate

B. Negligible

C. Slow

D. Very fast

Answer:

B. Negligible

viii) Which one of the following is a fat soluble vitamin:

(Mark 1)

A. A

B. E

C. K

D. All of these

Answer:

D. All of these

ix) You want to dry a gas, which one of the following salt you will use:

(Mark 1)

A. COCl₂

B. CaO

C. NaCl

D. NA₂SiO₃

Answer:

B. CaO

x) A reaction between an acid and a base produces: (Mark 1)

- A. Salt and water
- B. Salt and gas
- C. Salt and acid
- D. Salt and base

Answer:

- A. Salt and water

xi) Matte is a mixture of: (Mark 1)

- A. Cu_2S and FeS
- B. CuS and FeO
- C. Cu_2O and FeO
- D. FeS and CuS

Answer:

- A. Cu_2S and FeS

xii) In laboratory, urea was prepared by: (Mark 1)

- A. Berzelius
- B. Dalton
- C. Rutherford
- D. Wohler

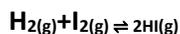
Answer:

- D. Wohler

Q.2 i) Write two properties of irreversible reactions. (Marks 2)

Q.2 ii) How atmospheric gases are used in manufacture of chemical? (Marks 2)

Q.2 iii) Write chemical equilibrium constant for given equation: (Marks 2)



Q.2 iv) What is complete reaction? How it is represented? (Marks 2)

Q.2 v) Write two properties of bases. (Marks 2)

Q.2 vi) According to Arrhenius, define acid, with an example. (Marks 2)

Q.2 vii) Write two properties of salts. (Marks 2)

Q.2 viii) Write the two uses of sodium carbonate. (Marks 2)

Q.3 i) Define functional group. (Marks 2)

Q.3 ii) Write classification of coal. (Marks 2)

- Q.3 iii) What are aromatic compounds? Give one example. (Marks 2)
- Q.3 iv) Why are alkenes reactive? (Marks 2)
- Q.3 v) Give two uses of ethene. (Marks 2)
- Q.3 vi) Give two characteristics of monosaccharides. (Marks 2)
- Q.3 vii) What is significance of vitamins? (Marks 2)
- Q.3 viii) How you justify RNA works as a messenger? (Marks 2)
- Q.4 i) Write the composition of dry air. (Marks 2)
- Q.4 ii) What are primary and secondary air pollutants? (Marks 2)
- Q.4 iii) Write two effects of ozone depletion. (Marks 2)
- Q.4 iv) Write two disadvantages of hard water. (Marks 2)
- Q.4 v) What is capillary action? (Marks 2)
- Q.4 vii) Write name of any two fractions of petroleum. (Marks 2)
- Q.4 viii) What role is played by pine oil in the froth flotation process? (Marks 2)
- Q.5 a) State the law of mass action and how chemical equilibrium constant is helpful in prediction of direction of reaction ? (Marks 5)
- Q.5 b) Write the uses of any four bases. (Marks 4)
- Q.6 a) Give five physical properties of alkenes. (Marks 5)
- Q.6 b) Explain the sources and uses of carbohydrates. (Marks 4)
- Q.7 a) How is urea manufactured? Explain with flow sheet diagram. (Marks 5)
- Q.7 b) How temporary hardness of water can be removed? explain (Marks 4)